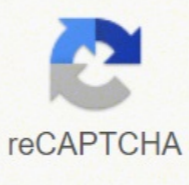




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Polyatomic Ions Worksheet Answer Key

Polyatomic Practice

1. Name or write the formula for the following polyatomic ions

sulfate	SO_4^{2-}	CO_3^{2-}	carbonate
nitrite	NO_2^-	MnO_4^{1-}	manganate
perphosphate	PO_4^{3-}	SO_5^{2-}	persulfate
hypoiodite	IO^-	BrO_2^{1-}	bromite
chlorite	ClO_2^-	CO_4^{2-}	percarbonate
phosphite	PO_3^{3-}	PO_5^{3-}	perphosphate
percarbonate	CO_4^{2-}	ClO^-	hypochlorite
bromate	BrO_3^-	IO_2^{1-}	iodite
hyposulfite	SO_2^{2-}	PO_4^{3-}	phosphate
permanganate	MnO_4^-	NO_2^{1-}	nitrite
carbonite	CO_2^{2-}	SO_4^{2-}	sulfate

2. Name or write the formula for the following Type I polyatomic ionic compounds

beryllium hydroxide	$Be(OH)_2$	$Ba(IO_3)_2$	barium iodate
sodium nitrite	$NaNO_2$	$Ga(CNO)_3$	gallium cyanate
ammonium chloride	NH_4Cl	Ag_2SO_3	silver sulfite
calcium bisulfate	$Ca(HSO_4)_2$	$MgCO$	magnesium hypocarbonite
rubidium perchlorate	$RbClO_4$	NH_3NO_2	ammonium nitrite
strontium sulfite	$SrSO_3$	$Al(C_2H_3O_2)_3$	aluminum acetate
aluminum acetate	$Al(C_2H_3O_2)_3$	$SrSO_4$	strontium persulfate
ammonium nitrate	NH_4NO_3	$RbClO_2$	rubidium chlorite
magnesium hypocarbonite	$MgCO$	$Ca(HSO_4)_2$	calcium bisulfate
silver hyposulfite	Ag_2SO_2	NH_4Cl	ammonium chloride

Ionic Chemical Formulas/PreAICE Chemistry

METAL	CATION	CATION FORMULA	NONMETAL ANION FORMULA	ANION FORMULA	COMPOUND NAME
sodium	Na+1	hydroxide	OH-	NaOH	sodium hydroxide
magnesium	Mg+2	phosphate	PO4-3	Mg3(PO4)2	magnesium phosphate
calcium	Ca+2	chloride	Cl-	CaCl2	calcium chloride
gold(I)	Au+1	sulfite	SO3-2	Au2SO3	gold(I) sulfite
copper(I)	Cu+1	hydroxide	OH-	CuOH	copper(I) hydroxide
iron(II)	Fe+2	telluride	Te-2	FeTe	iron(II) telluride
magnesium	Mg+2	bromide	Br-1	MgBr2	magnesium bromide
manganese(I)	Mn+1	sulfide	S-2	MnS	manganese(I) sulfide
boron	B+3	oxide	O-2	B2O3	boron oxide
potassium	K+	nitrate	NO3-	KNO3	potassium nitrate
barium	Ba+2	bromide	Br-	BaBr2	barium bromide
strontium	Sr+2	sulfate	SO4-2	SrSO4	strontium sulfate
iron(III)	Fe+3	oxide	O-2	Fe2O3	iron(III) oxide
silver(I)	Ag+1	chloride	Cl-	AgCl	silver(I) chloride
aluminum	Al+3	carbonate	CO3-2	Al2(CO3)3	aluminum carbonate
gold(III)	Au+3	nitride	N-3	AuN	gold(III) nitride
cesium	Cs+1	phosphate	PO4-3	Cs3PO4	cesium phosphate
ammonium	NH4+	sulfide	S-2	(NH4)2S	ammonium sulfide
zinc(II)	Zn+2	nitrite	NO2-	Zn(NO2)2	zinc(II) nitrite
chromium(III)	Cr+3	nitride	N-3	CrN	chromium(III) nitride
mercury(II)	Hg+2	oxide	O-2	HgO	mercury(II) oxide
sodium	Na+	sulfite	SO3-2	Na2SO3	sodium sulfite
magnesium	Mg+2	oxide	O-2	MgO	magnesium oxide

Name:

Date:

TOTAL: /40 Knowledge

Write the chemical formulas for the following. (20 marks)

a)	ammonium hypochlorite	NH_4ClO
b)	magnesium sulfate heptahydrate	$MgSO_4 \cdot 7H_2O$
c)	nitrous acid	$HNO_2(aq)$
d)	cesium tungstate	Cs_2WO_4
e)	bromine	$Br_2(l)$
f)	sodium sulfite	Na_2SO_3
g)	calcium carbide	Ca_2C / <i>Actually CaC₂</i>
h)	cupric acetate	$Cu(C_2H_3O_2)_2$
i)	sodium thiosulphite	$Na_2S_2O_2$
j)	iodine pentafluoride	I_2F_5
k)	barium molybdenate	$BaMoO_4$
l)	bismuthic selenide	Bi_2Se_5
m)	gallium chloride	$GaCl_3$
n)	aurous dihydrogenphosphite	AsH_2PO_2
o)	aluminum peroxide	Al_2O_4
p)	lead (IV) oxide	PbO_2
q)	nickelous fluorate	$Ni(FO_3)_2$
r)	lithium bicarbonite	$LiHCO_2$
s)	strontium phosphide	Sr_3P_2
t)	hydrocyanic acid	$HCN(aq)$

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